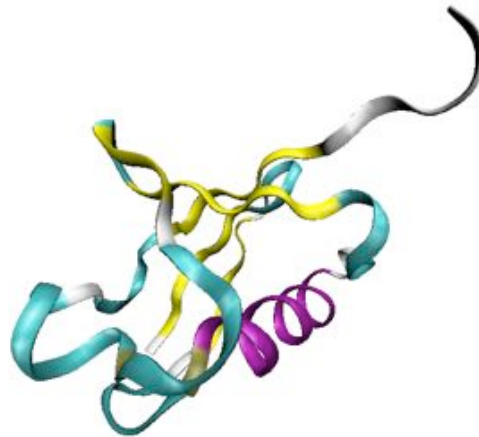


# Statistical Mechanics of Proteins



**Ubiquitin**

## Equilibrium Properties of Proteins

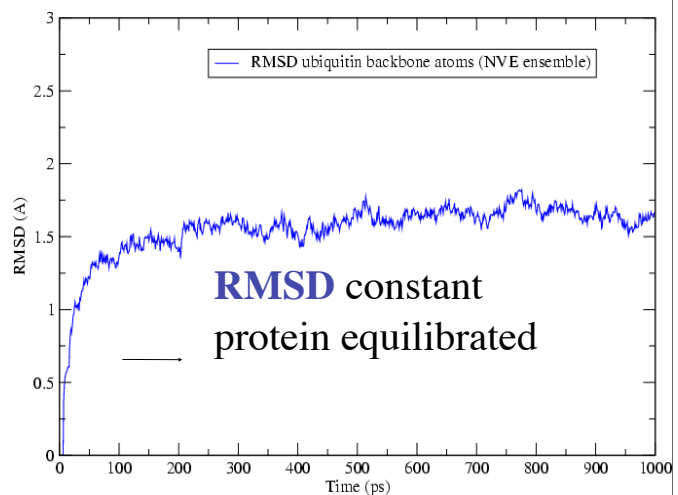
### Ubiquitin

Root Mean Squared Deviation: measure for equilibration and protein flexibility

$$RMSD_{\alpha} = \sqrt{\frac{\sum_{j=1}^{N_t} \sum_{\alpha=1}^{N_{\alpha}} (\vec{r}_{\alpha}(t_j) - \langle \vec{r}_{\alpha} \rangle)^2}{N_{\alpha}}}$$

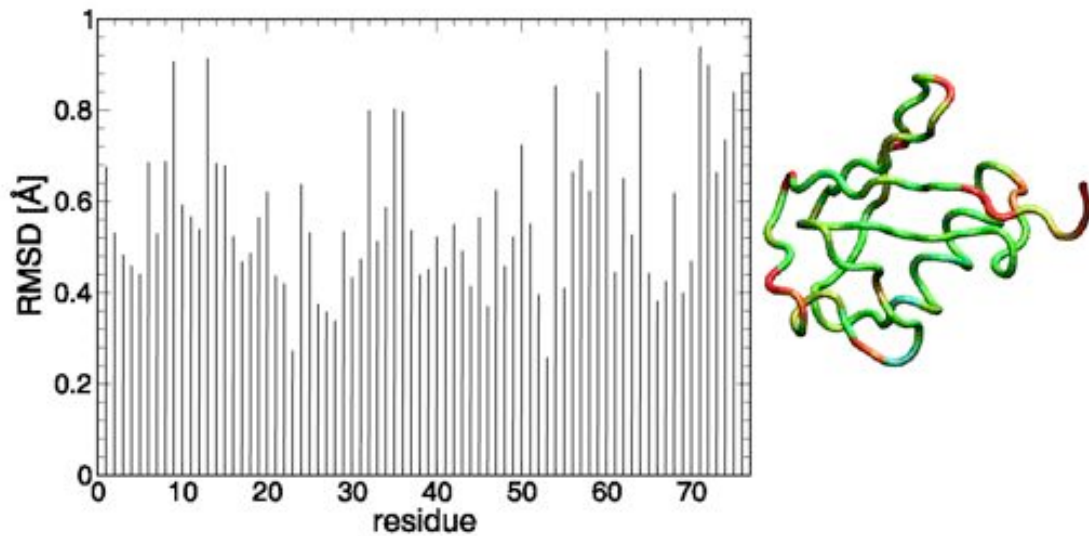


NMR structures  
aligned together to see flexibility



# Thermal Motion of Ubiquitin from MD

RMSD values per residue



**Protein sequence exhibits characteristic permanent flexibility!**

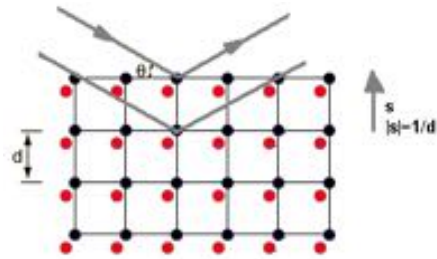
# Thermal Motion of Ubiquitin from MD

*Temperature Dependence of Crystal Diffraction (Debye-Waller factor)*

Bragg's law

$$2d \sin \theta = \lambda$$

due to contribution from each copy  $j$



$$f_j \exp[-i\vec{s} \cdot \vec{r}_j] \quad \text{structure factor } f_j$$

But the atom carries out thermal vibrations around equilibrium position  $\vec{x}_j$

$$\vec{r}_j(t) = \vec{x}_j + \vec{u}_j(t)$$

Accordingly:

$$\langle f_j \exp[-i\vec{s} \cdot \vec{r}_j] \rangle = f_j \exp[-i\vec{s} \cdot \vec{x}_j] \langle \exp[-i\vec{s} \cdot \vec{u}_j] \rangle$$

# Thermal Motion of Ubiquitin from MD

## Temperature Dependence of Crystal Diffraction (Debye-Waller factor)

One can expand:

$$\langle \exp[-i\vec{s} \cdot \vec{u}_j] \rangle = 1 - \underbrace{i \langle \vec{s} \cdot \vec{u}_j \rangle}_{=0} - \frac{1}{2} \langle (\vec{s} \cdot \vec{u}_j)^2 \rangle + \dots$$

Spatial average for harmonic oscillator:  $\langle (\vec{s} \cdot \vec{u}_j)^2 \rangle = \frac{1}{3} s^2 \langle u_j^2 \rangle$

One can carry out the expansion further and show all further terms vanish

$$\langle \exp[-i\vec{s} \cdot \vec{u}_j] \rangle = \exp \left[ -\frac{1}{6} s^2 \langle u_j^2 \rangle \right]$$

Using for the thermal amplitude of the harmonic oscillator

$$\frac{1}{2} m \omega^2 u_j^2 = \frac{3}{2} k_B T \quad \text{thermal "weakening"}$$

one obtains

Debye-Waller factor **of struct. factor**

$$\langle f_j \exp[-i\vec{s} \cdot \vec{r}_j] \rangle = f_j \exp[-s^2 k_B T / 2m\omega^2] \exp[-i\vec{s} \cdot \vec{x}_j]$$

# Equilibrium Properties of Proteins

Energies: kinetic and potential



temperature dependence

$$\begin{aligned}
 & \left\langle \sum_j \frac{1}{2} m_j v_j^2 \right\rangle && \text{Kinetic energy (quadratic)} \\
 U(\vec{R}) = & \underbrace{\sum_{\text{bonds}} k_i^{\text{bond}} (r_i - r_0)^2}_{U_{\text{bond}}} + \underbrace{\sum_{\text{angles}} k_i^{\text{angle}} (\theta_i - \theta_0)^2}_{U_{\text{angle}}} + \\
 & \underbrace{\sum_{\text{dihedrals}} k_i^{\text{dihe}} [1 + \cos(n_i \phi_i + \delta_i)]}_{U_{\text{dihedral}}} + \\
 & \underbrace{\sum_i \sum_{j \neq i} 4\epsilon_{ij} \left[ \left( \frac{\sigma_{ij}}{r_{ij}} \right)^{12} - \left( \frac{\sigma_{ij}}{r_{ij}} \right)^6 \right] + \sum_i \sum_{j \neq i} \frac{q_i q_j}{\epsilon r_{ij}}}_{U_{\text{nonbond}}}
 \end{aligned}$$

Potential energy (not all quadratic)

# Equilibrium Properties of Proteins

Energies: kinetic and potential



temperature dependence

$$\left\langle \sum_j \frac{1}{2} m_j v_j^2 \right\rangle = \frac{3}{2} N k_B T$$

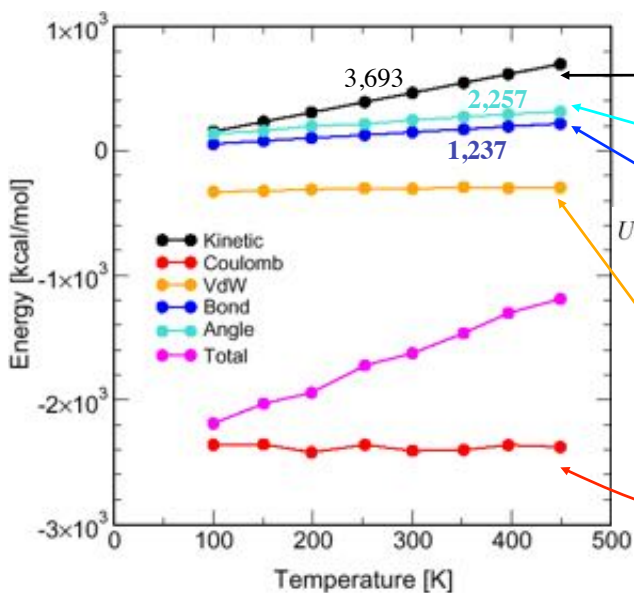
Kinetic energy (quadratic)

$$U(\vec{R}) = \underbrace{\sum_{\text{bonds}} k_i^{\text{bond}} (r_i - r_0)^2}_{U_{\text{bond}}} + \underbrace{\sum_{\text{angles}} k_i^{\text{angle}} (\theta_i - \theta_0)^2}_{U_{\text{angle}}} + \underbrace{\sum_{\text{dihedrals}} k_i^{\text{dihe}} [1 + \cos(n_i \phi_i + \delta_i)]}_{U_{\text{dihedral}}} + \underbrace{\sum_i \sum_{j \neq i} 4\epsilon_{ij} \left[ \left( \frac{\sigma_{ij}}{r_{ij}} \right)^{12} - \left( \frac{\sigma_{ij}}{r_{ij}} \right)^6 \right] + \sum_i \sum_{j \neq i} \frac{q_i q_j}{\epsilon r_{ij}}}_{U_{\text{nonbond}}}$$

Potential energy (not all quadratic)

# Equilibrium Properties of Proteins

Energies: kinetic and potential



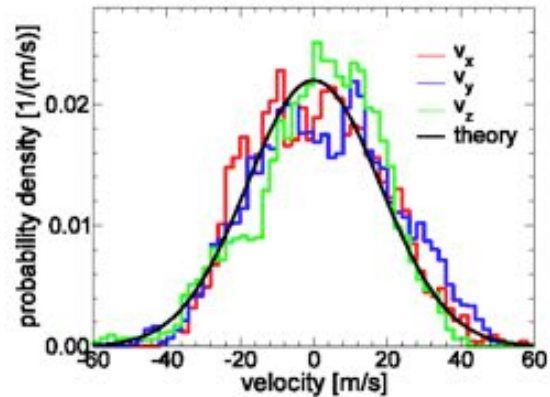
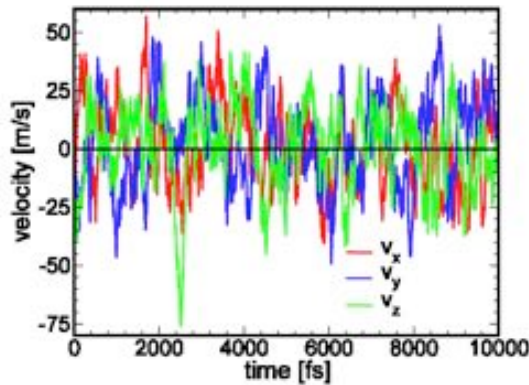
$$\left\langle \sum_j \frac{1}{2} m_j v_j^2 \right\rangle = \frac{3}{2} N k_B T$$

Kinetic energy (quadratic)

$$U(\vec{R}) = \underbrace{\sum_{\text{bonds}} k_i^{\text{bond}} (r_i - r_0)^2}_{U_{\text{bond}}} + \underbrace{\sum_{\text{angles}} k_i^{\text{angle}} (\theta_i - \theta_0)^2}_{U_{\text{angle}}} + \underbrace{\sum_{\text{dihedrals}} k_i^{\text{dihe}} [1 + \cos(n_i \phi_i + \delta_i)]}_{U_{\text{dihedral}}} + \underbrace{\sum_i \sum_{j \neq i} 4\epsilon_{ij} \left[ \left( \frac{\sigma_{ij}}{r_{ij}} \right)^{12} - \left( \frac{\sigma_{ij}}{r_{ij}} \right)^6 \right] + \sum_i \sum_{j \neq i} \frac{q_i q_j}{\epsilon r_{ij}}}_{U_{\text{nonbond}}}$$

Potential energy (not all quadratic)

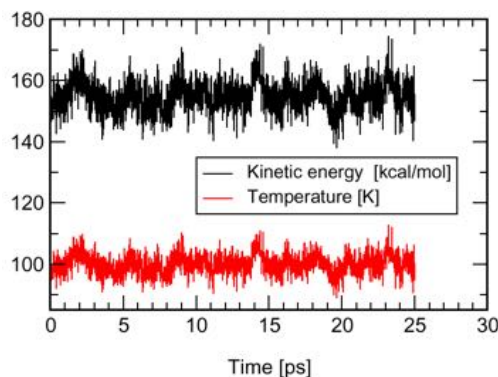
# Maxwell Distribution of Atomic Velocities



$$p(v_\sigma) = \sqrt{\frac{m}{2\pi k_B T}} \exp\left[-\frac{mv_\sigma^2}{2k_B T}\right]$$

$$\sigma = x, y, z$$

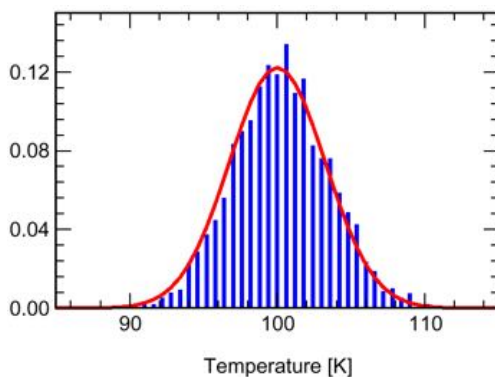
# Analysis of $E_{\text{kin}}$ , $T$ (free dynamics)



## Definition of Temperature

$$\left\langle \sum_j \frac{1}{2} m_j v_j^2 \right\rangle = \frac{3}{2} N k_B T$$

$$T = \frac{2}{3N k_B} \left\langle \sum_j \frac{1}{2} m_j v_j^2 \right\rangle$$



**The atomic velocities of a protein establish a thermometer, but is it accurate? Apparently not! Why!**

## Temperatur Fluctuations

Maxwell distribution

$$dP(v_n) = c \exp(-m v_n^2/2k_B T) dv_n \quad (7)$$

Individual kinetic energy  $\epsilon_n = m v_n^2/2$

$$dP(\epsilon_n) = (\pi T_0 \epsilon_n)^{-1/2} \exp(-\epsilon_n/k_B T_0) d\epsilon_n \quad (8)$$

One can derive

$$\langle \epsilon_n \rangle = T_0/2 \quad (9)$$

$$\langle \epsilon_n^2 \rangle = 3 T_0^2/4 \quad (10)$$

$$\langle \epsilon_n^2 \rangle - \langle \epsilon_n \rangle^2 = T_0^2/2 \quad (11)$$

The distribution of the total kinetic energy  $E_{kin} = \sum_j \frac{1}{2} m_j v_j^2$ , according to the central limit theorem, is approximately Gaussian

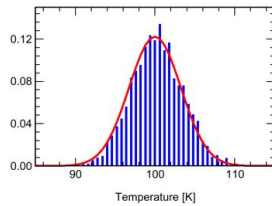
$$P(E_{kin}) = c \exp\left(\frac{-(E_{kin} - \langle E_{kin} \rangle)^2}{2 \left(\frac{3Nk_B^2 T_0^2}{2}\right)}\right) \quad (12)$$

The distribution function for the temperature ( $T = 2E_{kin}/3k_B$ ) fluctuations  $\Delta T = T - T_0$  is then

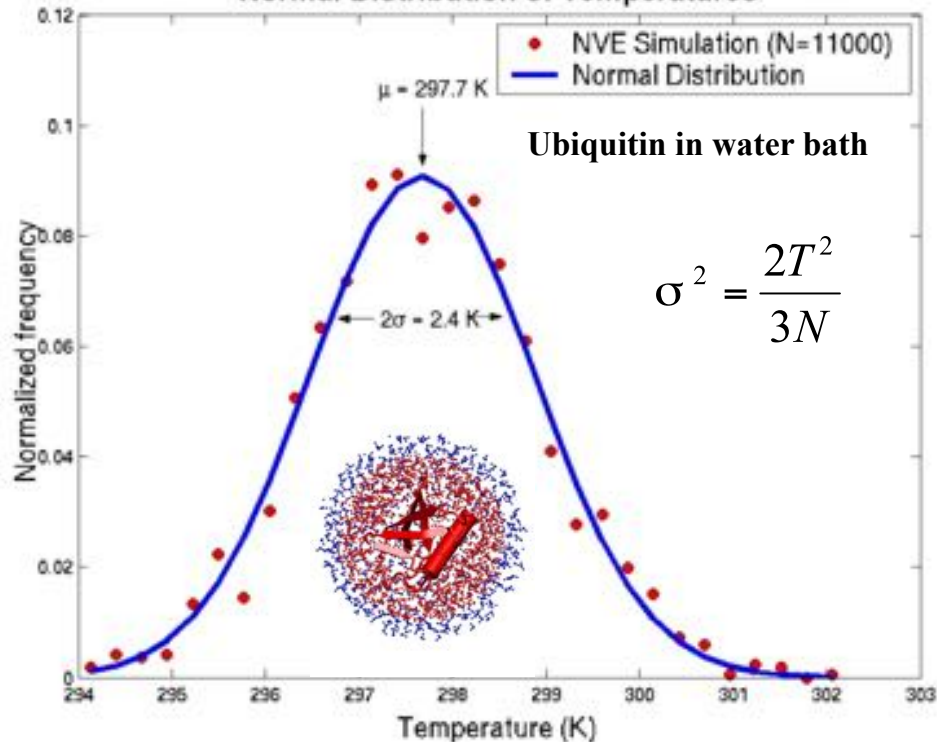
$$P(\Delta T) = c \exp[-(\Delta T)^2/2\sigma^2], \quad \sigma^2 = 2 T^2/3N \quad (13)$$

For  $T_0 = 100\text{K}$  and  $N = 557$ , this gives  $\sigma = 3.6$ .

The atomic velocity thermometer is inaccurate due to the finite size of a protein!



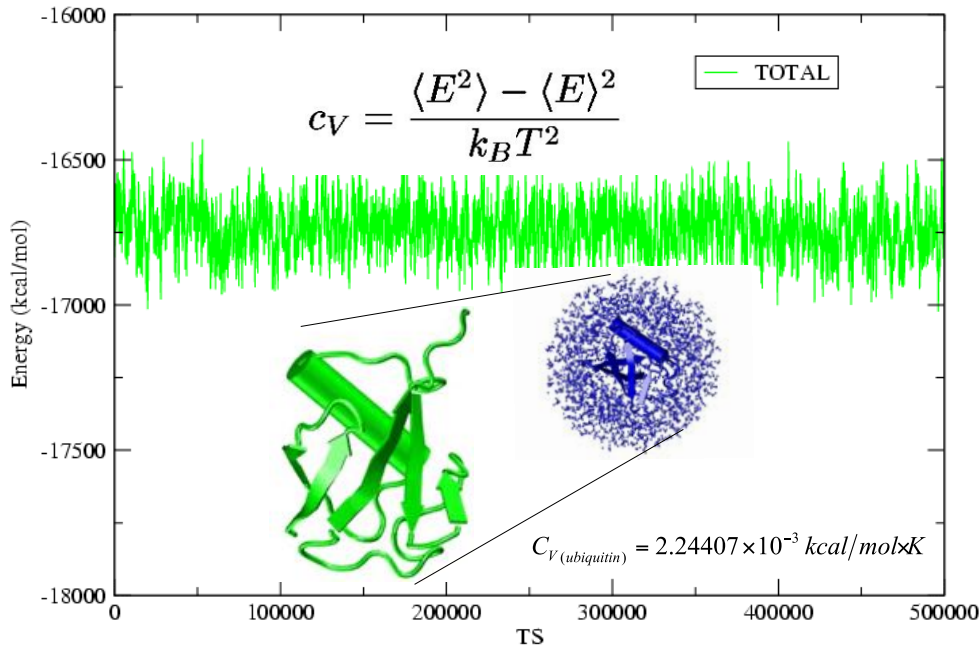
## Normal Distribution of Temperatures





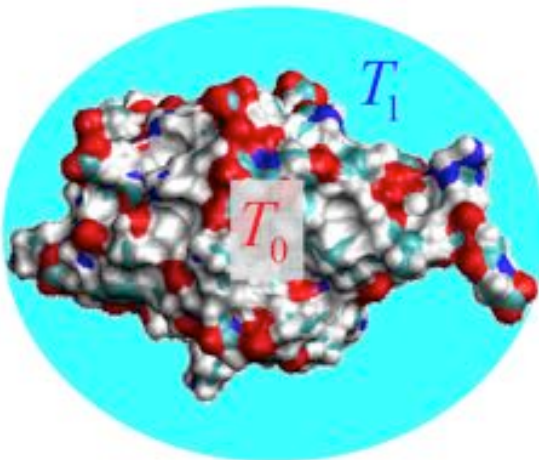
# Specific Heat of a Protein

Total energy of ubiquitin (NVE ensemble)



## Simulated Cooling of Ubiquitin

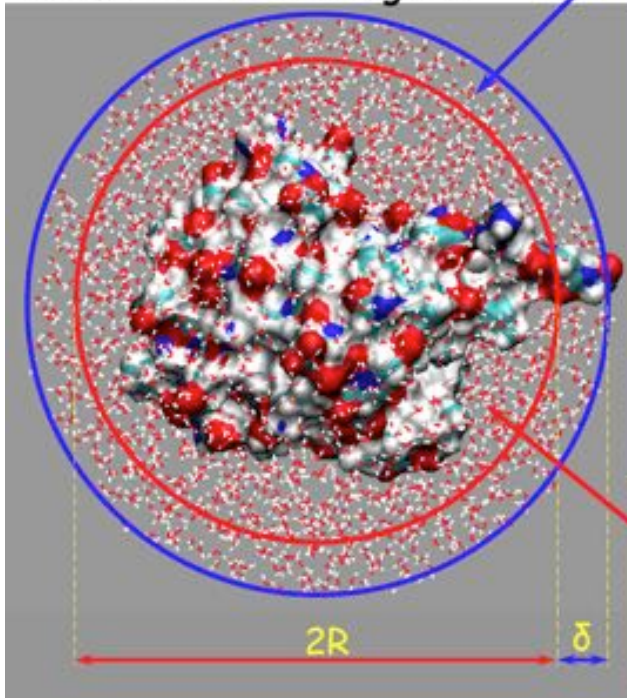
- Proteins function in a narrow (physiological) temperature range. What happens to them when the temperature of their surrounding changes significantly (temperature gradient) ?
- Can the heating/cooling process of a protein be simulated by molecular dynamics ? If yes, then how?



- What can we learn from the simulated cooling/heating of a protein ?

# How to simulate cooling ?

Heat transfer through mechanical coupling between atoms in the two regions



coolant layer of atoms

motion of atoms is subject to stochastic Langevin dynamics

$$m \ddot{\mathbf{r}} = \mathbf{F}_{FF} + \mathbf{F}_H + \mathbf{F}_f + \mathbf{F}_L$$

$\mathbf{F}_{FF}$  → force field

$\mathbf{F}_H$  → harmonic restrain

$\mathbf{F}_f$  → friction

$\mathbf{F}_L$  → Langevin force

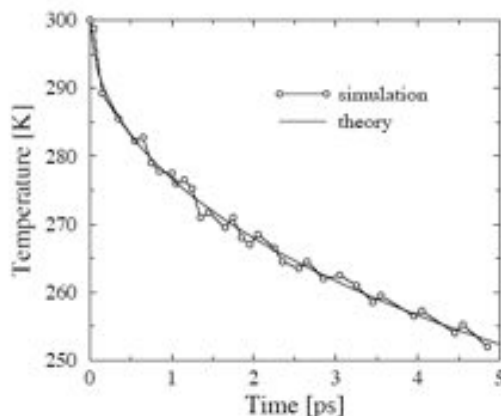
atoms in the inner region follow Newtonian dynamics

$$m \ddot{\mathbf{r}} = \mathbf{F}_{FF}$$

## Simulated Cooling - Result

$t$	$\langle T_{sim} \rangle$	$t$	$\langle T_{sim} \rangle$	$t$	$\langle T_{sim} \rangle$	$t$	$\langle T_{sim} \rangle$
0.05	298.75	1.05	276.00	1.95	267.00	3.25	261.00
0.15	289.25	1.15	276.50	2.05	268.50	3.45	258.50
0.35	285.50	1.25	275.25	2.25	266.50	3.55	259.50
0.55	282.25	1.35	271.00	2.35	264.50	3.95	256.50
0.65	282.75	1.45	271.75	2.55	263.50	4.05	257.25
0.75	279.00	1.65	269.50	2.65	264.50	4.45	254.00
0.85	277.75	1.75	271.00	2.85	262.00	4.55	255.25
1.00	277.50	1.85	268.00	3.05	262.50	4.85	252.00

Table 1: Mean temperature  $\langle T_{sim} \rangle$  [K] of the protein as a function of time  $t$  [ps].



Result from simulation that cools a protein, initially at  $T = 300$  K, from the edge down to 200 K. Only first 5 ps are shown in which short time interval  $T$  drops to 250 K.



## Heat Conduction Equation

$$\frac{\partial T(\mathbf{r}, t)}{\partial t} = D \nabla^2 T(\mathbf{r}, t)$$

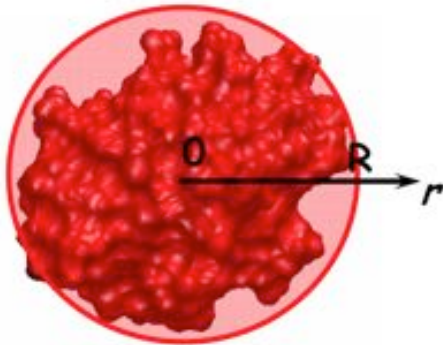
thermal diffusion  
coefficient

$$D = K / \rho c$$

mass density

specific heat

thermal conductivity



- approximate the protein with a homogeneous sphere of radius  $R \sim 20 \text{ \AA}$
- calculate  $T(r, t)$  assuming initial and boundary conditions:

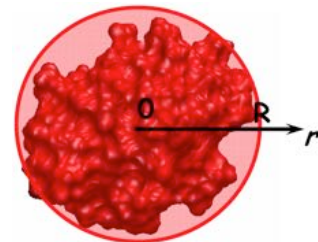
$$T(r, 0) = T_0 \text{ for } r < R$$

$$T(R, t) = T_{bath}$$

## Solution of the Heat Equation

$$\frac{\partial T(\mathbf{r}, t)}{\partial t} = D \nabla^2 T(\mathbf{r}, t),$$

$$D = K / \rho c,$$



**Initial condition**

$$T(\mathbf{r}, 0) = \langle T_{sim} \rangle(0) \text{ for } r < R,$$

**Boundary condition**

$$T(R, t) = T_{bath}.$$

# Solution of the Heat Equation

Spherical symmetry

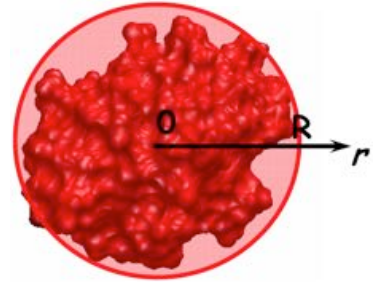
$$\frac{\partial T(r, t)}{\partial t} = D \frac{1}{r} \partial_r^2 r T(r, t)$$

$T_{bath}$

We assume

$$T(r, t) = T_{bath} + \sum_{n=1}^{\infty} a_n e^{\lambda_n t} u_n(r)$$

difference from bath



Here  $u_n$  are the eigenfunctions of the spherical diffusion operator

$$L \equiv \frac{D}{r} \frac{d^2}{dr^2} r$$

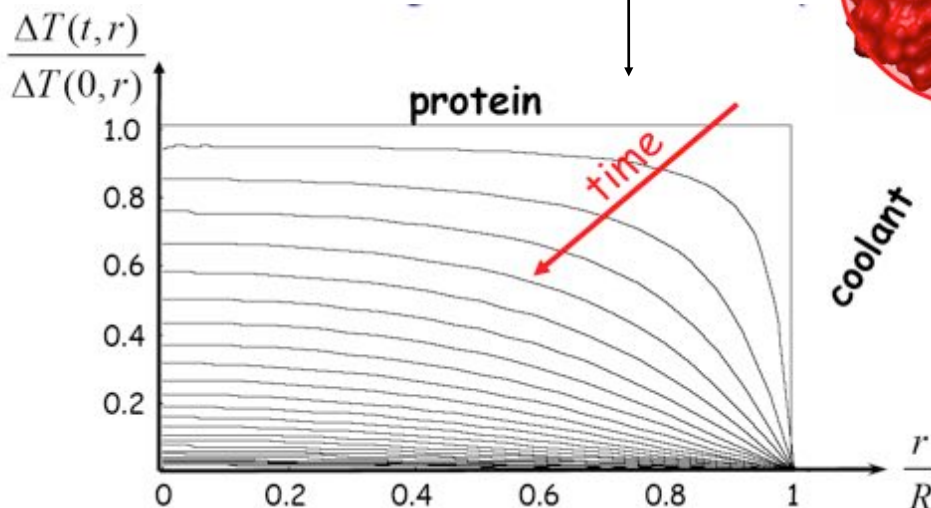
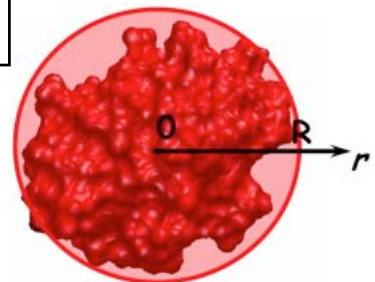
$$\frac{D}{r} \frac{d^2}{dr^2} r u_n(r) = \lambda_n u_n(r) \quad , \quad u_n(0) = \text{finite} \quad , \quad u_n(R) = 0$$

# Solution of the Heat Equation

$$T(r, t) = T_{bath} + \sum_{n=1}^{\infty} a_n \exp \left[ - \left( \frac{n\pi}{R} \right)^2 D t \right] \frac{\sin(n\pi r/R)}{r}$$

$$a_m = \frac{2R}{m\pi} \Delta T (-1)^{m+1}$$

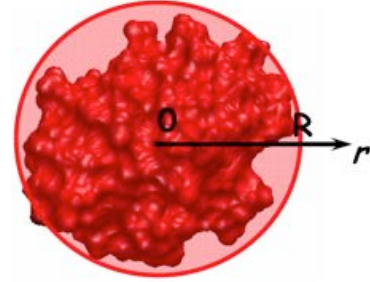
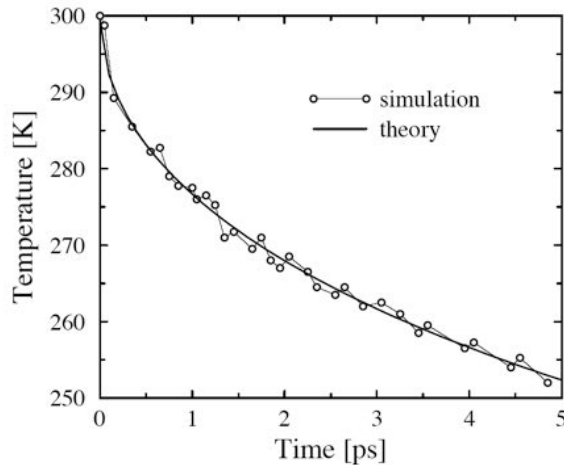
$T_{bath}$



# Solution of the Heat Equation

## Temperature averaged over volume

$$\begin{aligned}\langle T \rangle(t) &= \left( \frac{4\pi R^3}{3} \right)^{-1} \int d^3\mathbf{r} T(\mathbf{r}, t) = \frac{3}{R^3} \int_0^R r^2 dr T(r, t) \\ &= T_{bath} + \sum_{n=1}^{\infty} a_n \exp \left[ - \left( \frac{n\pi}{R} \right)^2 D t \right] \frac{3}{R^3} \int_0^R r dr \sin \left( \frac{n\pi r}{R} \right) \\ &= T_{bath} + 6 \frac{\Delta T}{\pi^2} \sum_{n=1}^{\infty} \frac{1}{n^2} \exp \left[ - \left( \frac{n\pi}{R} \right)^2 D t \right]\end{aligned}$$



$$D \approx 0.38 \times 10^{-3} \text{cm}^2 \text{s}^{-1}$$

$$\text{water } 1.4 \times 10^{-3} \text{cm}^2 \text{s}^{-1}$$